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Periodic Table With Chemistry Formulas SparkCharts

* Periodic Table with Chemistry Formulas

CHEMISTRY FORMULAS

ATOMS AND KINETIC PARTICLES

A. Atomic number: The number of protons in nucleus of an element is called atomic number. It is generally used to identify an element by writing its atomic number.

Ex: $^{12}_{6}\text{C}$ → 6 = atomic number

B. Atomic mass: The sum of mass of protons & neutrons in nucleus of an atom of an element is known as atomic mass. It is denoted by A . Unit of atomic mass is $1/12$ of ^{12}C atom.

Ex: $^{12}_{6}\text{C}$ → $12 = \text{atomic mass}$

Atomic number	Symbol	Mass	Charge
Hydrogen	^1_1H	1 amu	0
Helium	^4_2He	4 amu	0
Carbon	$^{12}_{6}\text{C}$	12 amu	0

C. Isotopes: Isotopes are atoms of an element that possess different numbers of neutrons, but have the same atomic number.

Ex: Isotopes of H : ^1_1H , ^2_1H , ^3_1H

D. Calculating atomic mass: An element's atomic mass is the weighted average of its isotopes.

Ex: $^{12}_{6}\text{C}$ → $12 = \text{atomic mass}$

E. Relative atomic mass: It is the ratio of the mass of an atom to $1/12$ of the mass of ^{12}C atom.

Ex: $^{12}_{6}\text{C}$ → $12 = \text{relative atomic mass}$

MOLECULAR MASS

A. Molecular mass: It is the sum of the atomic masses of all the atoms in a molecule.

Ex: H_2O → $2 \times 1 + 16 = 18$ amu

B. Molecular formula: It is the chemical formula of a compound that gives the actual number of atoms of each element in a molecule of the compound.

Ex: H_2O → Molecular formula

C. Empirical formula: It is the simplest formula of a compound that gives the ratio of the atoms in a molecule.

Ex: H_2O → Empirical formula

D. Molar mass: It is the mass of one mole of a substance.

Ex: H_2O → 18 g/mol

RELATIVE CHEMICAL EQUATIONS

A. Chemical equation: It is a chemical reaction that shows the reactants and products.

Ex: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

B. Balancing chemical equation: It is a process of making the number of atoms of each element equal on both sides of a chemical equation.

Ex: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

C. Stoichiometry: It is the study of the quantitative relationship between the reactants and products in a chemical reaction.

Ex: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

D. Limiting reagent: It is the reactant that is completely consumed in a chemical reaction.

Ex: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

E. Percent yield: It is the ratio of the actual yield to the theoretical yield, expressed as a percentage.

Ex: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

F. Percent composition: It is the percentage of each element in a compound, expressed as a percentage.

Ex: H_2O → $11.1\% \text{ H}$, $88.9\% \text{ O}$

G. Molar mass: It is the mass of one mole of a substance, expressed in grams per mole.

Ex: H_2O → 18 g/mol

H. Empirical formula: It is the simplest formula of a compound that gives the ratio of the atoms in a molecule.

Ex: H_2O → Empirical formula

I. Molecular formula: It is the chemical formula of a compound that gives the actual number of atoms of each element in a molecule of the compound.

Ex: H_2O → Molecular formula

J. Stoichiometry: It is the study of the quantitative relationship between the reactants and products in a chemical reaction.

Ex: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

K. Limiting reagent: It is the reactant that is completely consumed in a chemical reaction.

Ex: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

L. Percent yield: It is the ratio of the actual yield to the theoretical yield, expressed as a percentage.

Ex: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

M. Chemical equation: It is a chemical reaction that shows the reactants and products.

Ex: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

N. Balancing chemical equation: It is a process of making the number of atoms of each element equal on both sides of a chemical equation.

Ex: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

O. Stoichiometry: It is the study of the quantitative relationship between the reactants and products in a chemical reaction.

Ex: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

P. Limiting reagent: It is the reactant that is completely consumed in a chemical reaction.

Ex: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

Q. Percent yield: It is the ratio of the actual yield to the theoretical yield, expressed as a percentage.

Ex: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

R. Chemical equation: It is a chemical reaction that shows the reactants and products.

Ex: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$



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Synopsis

SparkCharts®: The information you need—concisely, conveniently, and accurately. Created by Harvard students for students everywhere, these study companions and reference tools cover a wide range of college and graduate school subjects, from Business and Computer Programming to Medicine, Law, and Languages. They'll give you what it takes to find success in school and beyond. Outlines and summaries cover key points, while diagrams and tables make difficult concepts easier to grasp. This two-page chart is a perfect reference for homework and problem sets. On side one, the chart lists the most important chemical formulas and provides quick refreshers on significant figures and balancing equations. Side two includes a beautiful periodic table that gives the following information for each element: Name Atomic number Atomic symbol Atomic mass Oxidation states (most stable state in bold) Electronegativity Electron affinity First ionization potential Atomic radius Electron shell configuration

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as advertised.

Very useful for my chem class at the university of Minnesota

A nice-looking periodic table for quick reference! What more could you want?

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